

QUESTION

In the following equilibrium:

 $HCO_3^{-}(aq) + H_2O(l) + H_2CO_3(aq) + OH^{-}(aq)$

- A) HCO_3 is an acid and H_2CO_3 is its conjugate base.
- B) H₂O is an acid and OH is its conjugate base.
- C) HCO₃ is an acid and OH is its conjugate base.
- D) H₂O is an acid and H₂CO₃ is its conjugate base.
- E) H₂O is an acid and HCO₃ is its conjugate base.

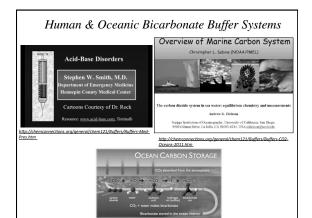
 $H_2CO_3(aq) / HCO_3^{-1}(aq) / CO_3^{-2}(aq)$

Two VERY IMPORTANT Buffer Systems

"Bicarbonate"

 $CO_2(g) + H_2O(l) \leftrightarrows HCO_3^{-1}(aq) + H^{+1}(aq) \leftrightarrows CO_3^{-2}(aq) + H^{+1}(aq)$

- Blood: a human's blood serum volume is relatively small, 4-6 Liters with a narrow pH range, pH = 7.35 – 7.45; pH is maintained through buffering (homeostasis) Have you ever had respiratory alkalosis during an exam?
- 2. Oceans: an extraordinarily large volume of a "salt water" solution with a pH \sim 8.1; maintained through buffering



EQUILIBRIUM

CO₂ & Oceanic Bicarbonate Buffering



 $CO_2(g) + H_2O(l) \leftrightarrows HCO_3^{-1}(aq) + H^{+1}(aq) \leftrightarrows CO_3^{-2}(aq) + H^{+1}(aq)$

Oceans: pH ~ 8.1 and falling

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